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Worksheet: Atomic Structure Name: _____

Use your notes from the Atomic Structure program to answer the following questions.

- The atomic number tells the number of positively charged _____ in the nucleus of an atom. The atom is _____ because this is also the number of _____ charged _____ in the atom.
- The mass number tells the total number of _____ and _____ in the nucleus of an atom. These particles collectively are called _____ since both are located in the nucleus.
- Isotopes are atoms of the same element with different numbers of _____ which results in different _____ numbers.
- Write the hyphen notation of the three isotopes of hydrogen:
 -
 -
 -
- Fill in the following chart (Hint: Isotopes may have different masses than the periodic table lists!)

Symbol	Atomic #	Mass #	# of protons	# of electrons	# of neutrons
Hg	80	200			
Bo	42	96			
	29			29	35
C		14			
	15	31			
Pb		208	82		
He	2				2

CHEMISTRY: A Study of Matter
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1-6

Atomic Theory and Structure of Matter

- Who proposed the atomic theory of matter?
- What is the basic unit of matter?
- What is the smallest particle of an element that cannot be created or destroyed in a chemical reaction?
- What is the smallest particle of a substance that cannot be created or destroyed in a chemical reaction?
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Name: _____ Date: _____

**Science 10: Chemistry
Atoms and The Periodic Table Worksheet**

- Define chemistry.
- Define atom.
- The three parts of an atom and their respective charges are :
- The atomic number of an atom tells you the number of _____
If the atom is neutral the atomic number also gives the number of _____
No two elements will have the _____ atomic number.
- The atomic mass of an atom can be used to determine the number of _____
First _____ the atomic mass and then complete the following equation:
of Neutrons = _____ - _____
Use your periodic table to complete the following:
- Give the symbol and number of protons in one atom of:

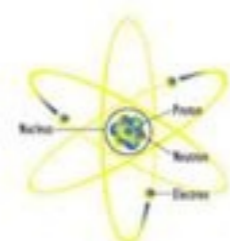
Lithium	_____	Bromine	_____
Iron	_____	Copper	_____
Oxygen	_____	Mercury	_____
Krypton	_____	Helium	_____
- Give the symbol and number of electrons in a neutral atom of:

Uranium	_____	Chlorine	_____
Boron	_____	Iodine	_____
Antimony	_____	Xenon	_____
- Give the symbol and number of neutrons in one atom of:

Barium	_____	Bismuth	_____
Carbon	_____	Hydrogen	_____
Fluorine	_____	Magnesium	_____
Europium	_____	Mercury	_____
- Name the element which has the following number of particles:

26 electrons, 29 neutrons, 26 protons	_____
53 protons, 74 neutrons	_____
2 electrons (neutral atom)	_____
20 protons	_____
0 neutrons	_____
- If you know only the following information, can you always determine what the element is? Yes or no?

Number of protons	_____
Number of neutrons	_____
Number of electrons in a neutral atom	_____
Number of electrons	_____



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